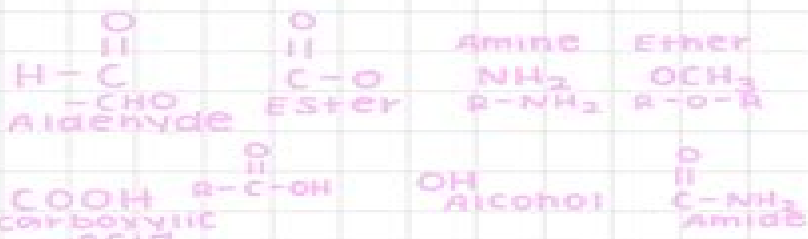
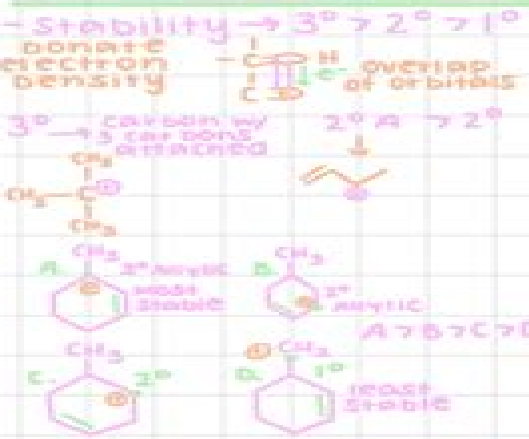


# FUNCTIONAL GROUPS



# CARBOCATIONS



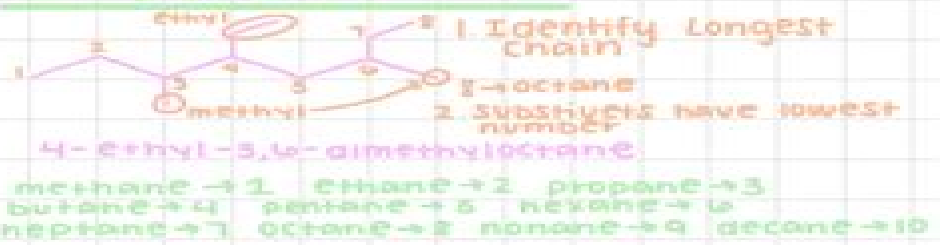
# LEWIS BASE

- nucleophile  $\rightarrow$  electron rich
- can donate electrons
- negative charge
- $\text{HO}^- \rightarrow$  hydroxide
- radical cannot act as LB  $\rightarrow \text{CH}_3^\bullet$
- need lone pair to donate
- $\text{H}_2\text{O}$  acts as Lewis base

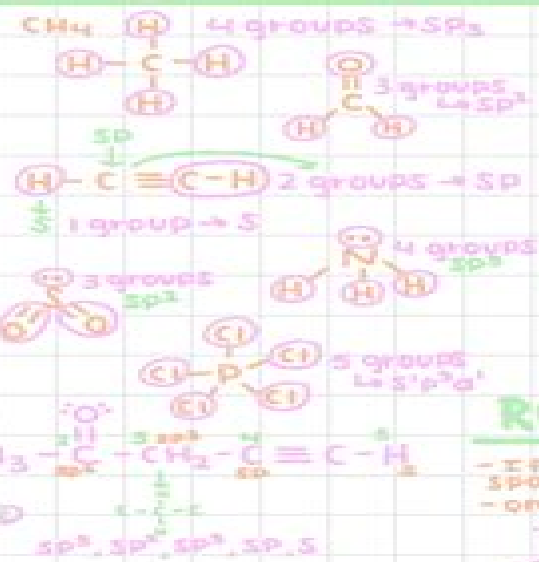
# CONFIGURATION

- Rank each group based on atomic numbers
  - 5 > Br > F > CH<sub>3</sub> ignore H
  - clockwise  $\rightarrow R$
  - counterclockwise  $\rightarrow S$
  - H in the back
- 1 COOH  
2 CH<sub>2</sub>OH  
3 CH<sub>2</sub>  
4 H
- since H is in front so reverse

# IUPAC NAME



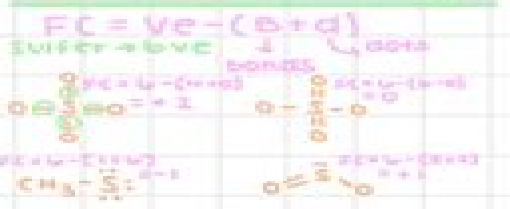
# HYBRIDIZATION



# ACIDS

- $\text{pKa} \downarrow$  strength  $\uparrow$
- $\text{CH}_3\text{COOH}$   $\text{CH}_3\text{CH}_2\text{OH}$   
carboxylic acid  $\text{pKa} \approx 4$  alcohol  $\text{pKa} \approx 16$
- $\text{CH}_3\text{NH}_2$   
amine  $\text{pKa} \approx 35-40$
- $\text{CH}_3\text{SO}_3\text{H}$   $\text{pKa} \approx 1$   
sulfonic acid  $\rightarrow$  very strong
- if you don't know pKa
- compare element connected to H
  - O > N > C
  - F > Cl > Br > I
  - $\text{H}_2\text{O} > \text{NH}_3 > \text{CH}_4$

# FORMAL CHARGE



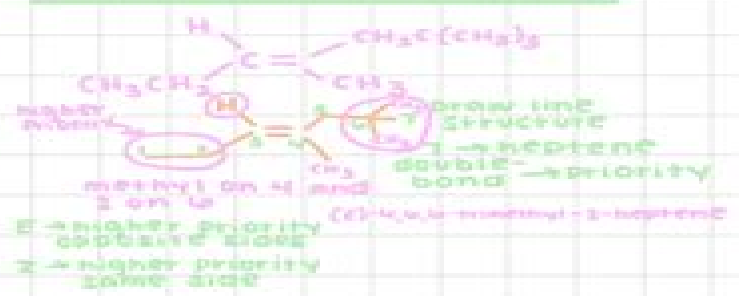
# LEWIS STRUCTURE



# RESONANCE STRUCTURE

- If any atom is moving from one spot to another it is NOT a RS
  - only electron is free to move
- only  $e^-$  move

# IUPAC NAME



# Organic Chemistry 1 Final Exam Cheat Sheet

**Michael Seilmaier**



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