

Functional Groups



Carbocations

- Stability $\rightarrow 3^\circ > 2^\circ > 1^\circ$
 donate electron density
 overlap of orbitals



A > B > C > D



least stable

Lewis Base

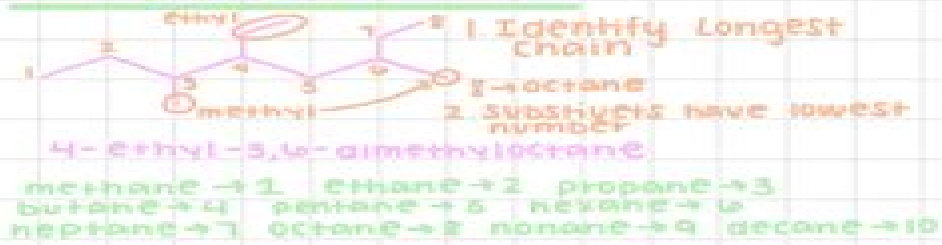
- Nucleophile \rightarrow electron rich
- can donate electrons
- negative charge
- $\text{HO}^- \rightarrow$ hydroxide
- radical cannot act as LB $\rightarrow \text{CH}_3^\bullet$
- need lone pair to donate
- H_2O acts as Lewis base

Configuration

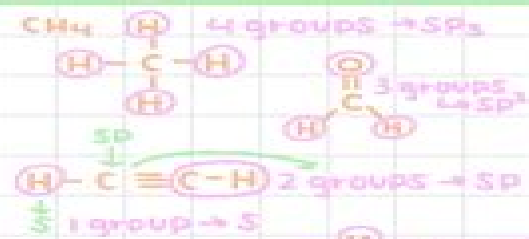
- Rank each group based on atomic numbers
- 5 > Br > F > CH₃ ignore H
- clockwise $\rightarrow R$
- counterclockwise $\rightarrow S$
- H in the back



IUPAC Name



Hybridization



Acids

$\text{pK}_a \downarrow$ strength \uparrow

CH_3COOH $\text{CH}_3\text{CH}_2\text{OH}$
 carboxylic acid $\text{pK}_a \approx 4-5$ alcohol $\text{pK}_a \approx 15-16$

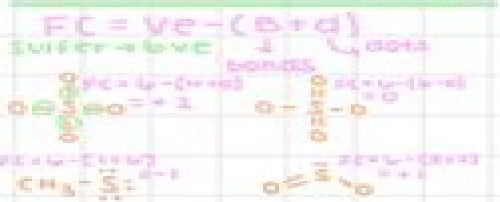
CH_3NH_2 CH_3COOH
 amine $\text{pK}_a \approx 35-40$ carboxylic acid $\text{pK}_a \approx 4-5$

$\text{CH}_3\text{SO}_3\text{H}$ CH_3COOH
 sulfonic acid $\text{pK}_a \approx -1$ carboxylic acid $\text{pK}_a \approx 4-5$

If you don't know pK_a

- compare element connected to H
- $\text{OH} > \text{NH}_2 > \text{C} > \text{N} > \text{O} > \text{F}$
- $\text{HI} > \text{HBr} > \text{HCl} > \text{HF}$
- $\text{H}_2\text{O} > \text{NH}_3 > \text{CH}_4$

Formal Charge



Lewis Structure

Cyanide ion (CN^-)

$\text{C} \equiv \text{N}^-$

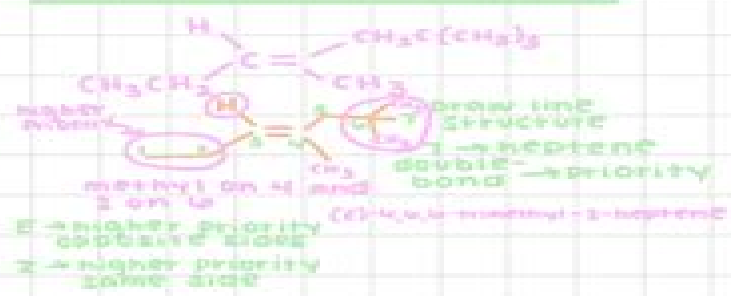
- must follow octet rule
- most stable \rightarrow minimal FC

Resonance Structure

- If any atom is moving from one spot to another it is NOT a RS
- only electron is free to move

$\text{O}=\text{C}=\text{O} \leftrightarrow \text{O}^+=\text{C}=\text{O}^-$ only e^- move

IUPAC Name



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J Dewey



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